

## TOPOLOGICAL RELATIONS IN THE SYSTEM U-Si-O-H FORMED DURING THE WEATHERING OF THE ALBITITE-TYPE URANIUM CACHOEIRA DEPOSIT, CAETITÉ, BAHIA, BRAZIL

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**RESUMO:** The alteration pattern at the Cachoeira Mine due to weathering results in the formation of uranophane  $[\text{Ca}(\text{UO}_2)_2(\text{SiO}_3)_2(\text{OH})_2]$ , soddyite  $[\text{UO}_2)_5(\text{SiO}_4)_2(\text{OH})_2 \cdot 5\text{H}_2\text{O}]$ , kasolite  $\text{Pb}[\text{UO}_2)(\text{SiO}_3)(\text{OH})_2]$  with very subordinate amounts of coffinite  $[\text{USiO}_4)_{4/0.5}(\text{OH})_2]$ , formed under high silica activity. Hydration of primary uraninite or secondary pitchblende, which could be partly destabilized during weathering, under silica- and/or water-rich conditions form the encountered mineral association. Nine key reactions have been assembled in  $\mu\text{SiO}_2$ - $\mu\text{H}_2\text{O}$  diagram: uraninite +  $\frac{1}{2}\text{SiO}_2 + \text{H}_2\text{O} = \text{coffinite}$  (1); 2 uraninite +  $2\text{SiO}_2 + 6\text{H}_2\text{O} = \text{uranophane}$  (2); uraninite +  $\text{SiO}_2 + \text{H}_2\text{O} = \text{kasolite}$  (3); 5 uraninite +  $2\text{SiO}_2 + 6\text{H}_2\text{O} = \text{soddyite}$  (4); uranophane = 2 soddyite +  $4\text{H}_2\text{O}$  (5); uranophane = 2 coffinite +  $\text{SiO}_2 + 4\text{H}_2\text{O}$  (6); soddyite +  $3\text{SiO}_2 = 5\text{kasolite} + \text{H}_2\text{O}$  (7); 4 coffinite +  $\text{H}_2\text{O} = \text{soddyite} + \text{SiO}_2$  (8); coffinite +  $\frac{1}{2}\text{SiO}_2 = \text{kasolite}$  (9). It could be shown that increasing  $\mu\text{H}_2\text{O}$  would make stable uranophane. Under silica-rich conditions kasolite and uranophane form and moderate  $\mu\text{SiO}_2$ - $\mu\text{H}_2\text{O}$  soddyite and coffinite are the stable phases. Although it is difficult to detect and recognize other phases such as becquerelite  $[\text{Ca}(\text{UO}_2)_6\text{O}_4(\text{OH})_6(\text{H}_2\text{O})_8]$ , sklodowskite  $[\text{Mg}[\text{UO}_2)_2(\text{SiO}_3)_2(\text{OH})_2 \cdot 5\text{H}_2\text{O}]$  and schoepite  $[\text{UO}_2)_8\text{O}_2(\text{OH})_{12} \cdot (\text{H}_2\text{O})_{12}]$  because they are poorly crystallized or amorphous, eventually they could exist and included in the reaction set, enlarging the number of reactions to twenty one. The uranyl oxyhydroxides form in U-rich aqueous condition through oxidation and hydration of uraninite. In that way, all eight mineral phases may be represented in a single diagram. Increasing  $\mu\text{H}_2\text{O}$  favor stability of becquerelite, which under still higher  $\text{H}_2\text{O}$  conditions yield to schoepite. Increasing both  $\text{H}_2\text{O}$  and silica, sklodowskite must form. As these minerals still have not been positively found at Cachoeira, we assume that high silica/ $\text{H}_2\text{O}$  ratio prevailed

**PALAVRAS-CHAVE:** TOPOLOGY; U-SECONDARY MINERALS; MINE 13-CACHOEIRA.